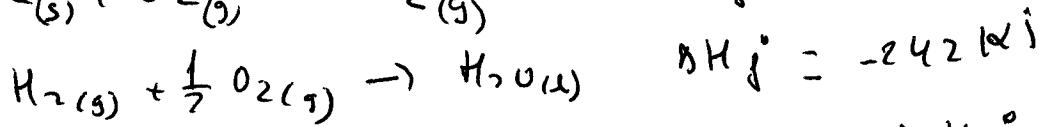
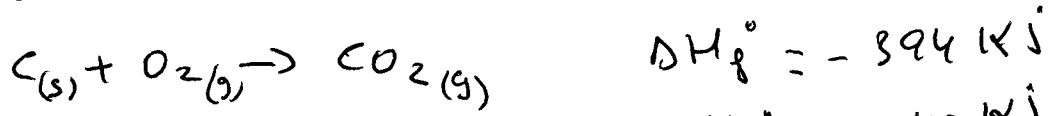
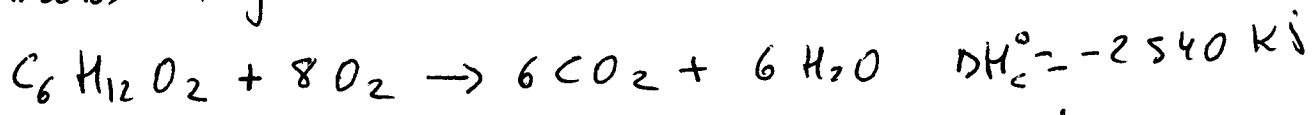


④ Datos en forma de ecuaciones termoquímicas:

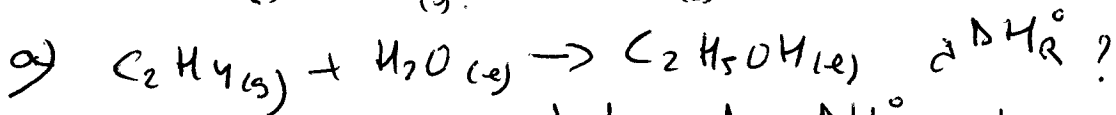
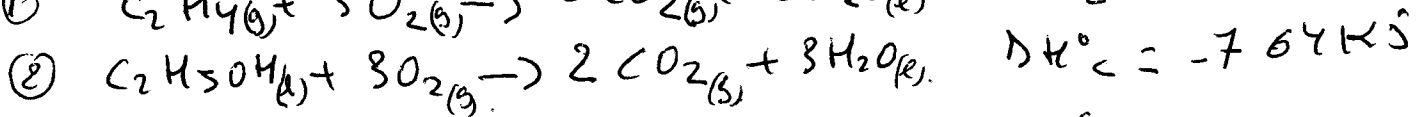
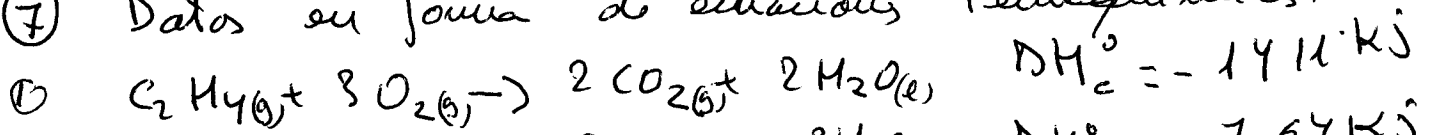


Aplicando  $\Delta H_R = \sum \Delta H_P - \sum \Delta H_R$  a la ec. 1º

$$-2540 = 6 \cdot (-394) + 6(-242) - (X + 0)$$

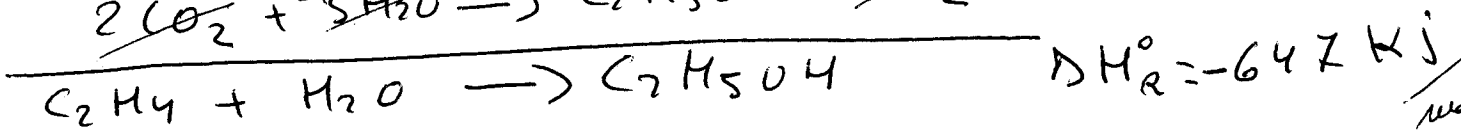
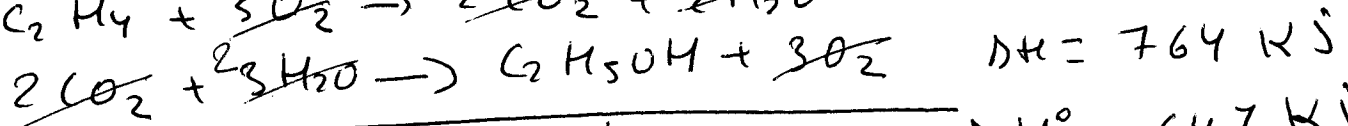
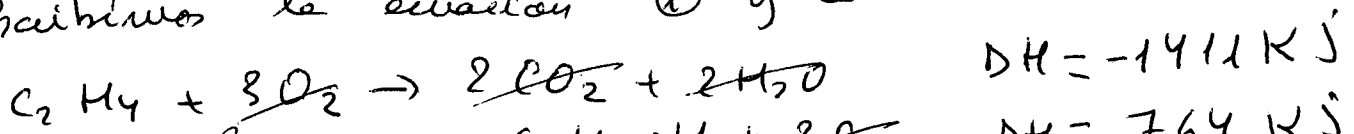
Despejamos  $X \Rightarrow X = \Delta H_f^\circ C_6H_{12}O_2 = -1276 \text{ KJ/mol}$

⑦ Datos en forma de ecuaciones termoquímicas:



No tenemos como dato la  $\Delta H_f^\circ$ , luego debemos resolver por la ley de Hess.

Escribimos la ecuación ① y la inversa de la ② =



b)  $P.M. C_2H_5OH = 46 \text{ g/mol}$

$$75 \text{ g } C_2H_5OH \cdot \frac{1 \text{ mol}}{46 \text{ g}} \cdot \frac{647 \text{ KJ}}{1 \text{ mol}} = 1054 \text{ KJ}$$

Como  $1054 \text{ KJ} \Rightarrow \Delta H = -1054 \text{ KJ}$